

Unit 8 CYOA Quiz

Name _____ Date _____

_____ **5.5**

Formula

Sodium Oxide

Iron (III) Nitride

Magnesium Carbonate

_____ **5.6**

Formula

Dinitrogen Pentoxide

Sulfur Dioxide

Carbon Tetrachloride

_____ **7.1**

How many moles are in 75.6 g of sodium?

_____ **7.2**

How many atoms are in 4.75 moles of copper?

How many grams are in 5.3 moles of potassium?

How many moles are in 7.2×10^{24} atoms of beryllium?

_____ **7.3**

You mix 10.0 g of HCl in 500. mL of water. What is the molarity?

You have 450. L of a 0.500 M solution of NaCl. How many moles of NaCl do you have?

_____ **8.1** Sodium reacts with chlorine gas to form sodium chloride. If you have 8.0 moles of chlorine gas, how many moles of sodium chloride can you make?

_____ **8.2 (must complete w/ 8.1)**

How many grams of sodium chloride is this?

_____ **8.3 (must complete w/8.1 & 8.2)**

If you produce 900. g of NaCl, what is the % yield?

_____ **8.4** Sodium reacts with chlorine gas to form sodium chloride. If you have 8.0 moles of chlorine gas and 10.0 moles of sodium, how many moles of sodium chloride can you make?

_____ **8.5 (must complete w/ 8.4)** How many liters of Cl_2 do you need to use in the problem above if the reaction occurs at STP?