

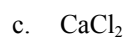
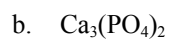
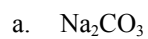
## UNIT -- Quantities in Chemical Reactions

### LESSON -- Moles and Molar Mass of Compounds

Name: \_\_\_\_\_

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1. Determine the molar mass of the following compounds:



2. A pop can contains around  $3.34 \times 10^{23}$  aluminum atoms. How many moles of aluminum are contained in a pop can?

3. The average glass of water contains 235.8 g of water molecules. How many moles of water molecules are in the average glass of water?
4. Iron (Fe) is a common building material. A standard 2-inch finishing nail contains about  $8.55 \times 10^{21}$  iron atoms. What is the mass of a finishing nail?
5. An underground cavern is completely filled with 8198 g of methane gas ( $\text{CH}_4$ ). What is the volume (in L) of the underground cavern? (Assume the gas is at STP conditions)