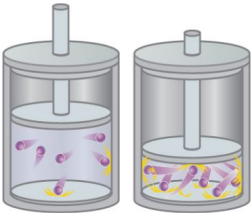
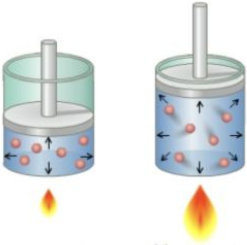
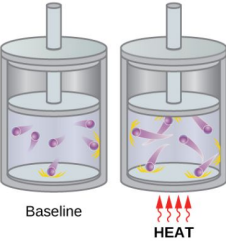
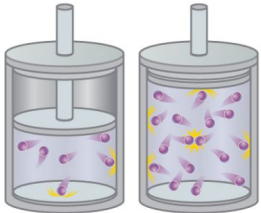
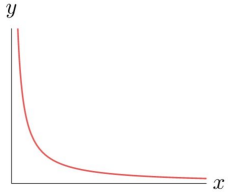
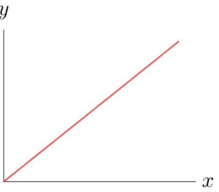
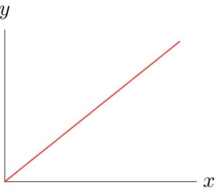
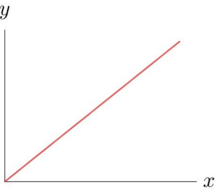


Boyle's Law (P-V)	Charles' Law (V-T)	Gay-Lussac's Law (P-T)	Avagadro's Law (n-V)
$P_1V_1 = P_2V_2$	$\frac{V_1}{T_1} = \frac{V_2}{T_2}$	$\frac{P_1}{T_1} = \frac{P_2}{T_2}$	$\frac{n_1}{V_1} = \frac{n_2}{V_2}$
Decreasing volume decreases the space particles have to move which increases particle collisions, increasing pressure.	Increasing temperature increases the energy of the particles. Faster particles collide harder and push each other farther apart, increasing volume.	Increasing temperature increases the energy of the particles. Particles collide harder and more often against a rigid container, increasing pressure.	Increasing the number of molecules in a container causes the particles to collide more often and push each other farther apart, increasing volume.
			
			
A 5.0 L sample of a gas is at 1.5 atm. The pressure is lowered to 1.0 atm. What is the new volume?	A 5.00 L sample of a gas is at 200. K. The temperature is raised to 500. K. What is the new volume?	A sample of a gas is at 2.0 atm and 100 K. The temperature is raised to 500 K. What is the new pressure?	A 5.0 L sample of gas contains 5.0 moles. 10. more moles are added. What is the new volume?
Marshmallow in syringe	Heat flask with balloon	Can crush	Blow up a balloon

<b>1</b>	<b>2</b>	<b>3</b>	<b>4</b>
1	2	3	4
1	2	3	4
1	2	3	4
1/2/3	1/2/3	1/2/3	4
1 (15 L)	2 (10 L)	3 (12.5 L)	4 (7.5 L)
1	2	3	4

