

Polyatomic Ions = "many atoms" ions – ions made of _____

_____ bonded together

→ Treat them as _____

Name	Formula
-3 Charge	
Phosphate	PO_4^{3-}
-2 Charge	
Carbonate	CO_3^{2-}
Chromate	CrO_4^{2-}
Dichromate	$Cr_2O_7^{2-}$
Hydrogen phosphate	HPO_4^{2-}
Peroxide	O_2^{2-}
Sulfate	SO_4^{2-}
Sulfite	SO_3^{2-}
-1 Charge	
Acetate	$C_2H_3O_2^-$
Chlorate	ClO_3^-
Chlorite	ClO_2^-
Cyanide	CN^-
Dihydrogen phosphate	$H_2PO_4^-$
Hydrogen carbonate (bicarbonate)	HCO_3^-
Hydrogen sulfate	HSO_4^-
Hydroxide	OH^-
Hypochlorite	ClO^-
Nitrate	NO_3^-
Nitrite	NO_2^-
Perchlorate	ClO_4^-
Permanganate	MnO_4^-
Superoxide	O_2^-
+1 Charge	
Ammonium	NH_4^+
Hydronium	H_3O^+

****MEMORIZE THE BLUE AND GREEN HIGHLIGHTED IONS****

Polyatomic Names:

_____ = more O

_____ = less O

_____ = less than

_____ = more than

In Compounds:

-Treat them just as monatomic

ions are treated

-If you need more than 1 of the

ion to create a neutral

compound, use _____ around it

with the subscript

Examples:

*NaOH :

*magnesium acetate:

*ammonium sulfate: