

\* Name:

Example: *Oxide*



Example: *aluminum*



\* A \_\_\_\_\_ is a \_\_\_\_\_ charged ion

→ formed by the \_\_\_\_\_ of electrons

- \_\_\_\_\_ form cations

-atoms with \_\_\_\_\_ valence electrons form \_\_\_\_\_

\* Name:

Example: *chloride*



Example: *nitride*



\* A \_\_\_\_\_ is a \_\_\_\_\_ charged ion

→ formed by the \_\_\_\_\_ of electrons

- \_\_\_\_\_ form anions

-atoms with \_\_\_\_\_ valence electrons form \_\_\_\_\_

\* Name:

Example: *Oxide*



Example: *aluminum*



\* A \_\_\_\_\_ is a \_\_\_\_\_ charged ion

→ formed by the \_\_\_\_\_ of electrons

- \_\_\_\_\_ form cations

-atoms with \_\_\_\_\_ valence electrons form \_\_\_\_\_

\* Name:

Example: *chloride*



Example: *nitride*



\* A \_\_\_\_\_ is a \_\_\_\_\_ charged ion

→ formed by the \_\_\_\_\_ of electrons

- \_\_\_\_\_ form anions

-atoms with \_\_\_\_\_ valence electrons form \_\_\_\_\_

# ALL ABOUT IONS

Charges:

Valence electrons:

Electrons found in the

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→ Called the

A periodic table of elements showing the number of valence electrons for each element. The valence electrons are indicated by a bracket on the right side of the table, encompassing the s and p orbitals for the main group elements. The table includes elements from Hydrogen (H) to Oganesson (Og), with the lanthanide and actinide series shown below the main body.

**OCTET RULE:** Elements want to achieve a full octet:

Exceptions: \_\_\_\_\_ and \_\_\_\_\_ follow the duet rule, because...

Paste Cations  
and Anions  
Foldable Here