

<p>E°_{cell} is:</p> <p>If a cell reaches equilibrium, then...</p>	<p>The greater the value of E°_{cell}, the</p> <p>...the cell is</p>	<p>Meaning of E_{cell}</p>
<p>When $\Delta G < 0...$</p> <p>When $\Delta G > 0...$</p> <p>When $\Delta G = 0...$</p>	<p>...E°_{cell} is \rightarrow _____ cell</p> <p>...E°_{cell} is \rightarrow _____ cell</p> <p>...E°_{cell} is \rightarrow _____ cell</p>	<p>Relationship between ΔG and E°_{cell}</p>
<p>E° corresponds to standard cell conditions when</p>	<p>Standard Conditions: Concentration = _____ Temperature = _____</p>	<p>E°_{cell} for Standard Conditions</p>
<p>The Nernst Equation helps us understand electrochemical cells at nonstandard conditions:</p> <p>$E = E^\circ - (RT/nF) \ln Q$</p> <p>Nonstandard conditions allow us to create concentration cells</p> <p>* You will NOT need to do calculations with the Nernst Equation. You will only have to use this equation to QUALITATIVELY discuss changes of concentration.</p>	<p>\rightarrow If Q is less than 1: \rightarrow then the cell is _____ equilibrium \rightarrow the cell potential will _____ relative to E°.</p> <p>\rightarrow If Q is greater than 1: \rightarrow then the cell is _____ equilibrium \rightarrow the cell potential will _____ relative to E°.</p>	<p>E_{cell} for Nonstandard Conditions</p>
<p>DO use arguments that rely on:</p>	<p>DO NOT use arguments that rely on:</p>	<p>Discussing changes in electrochemical systems</p>