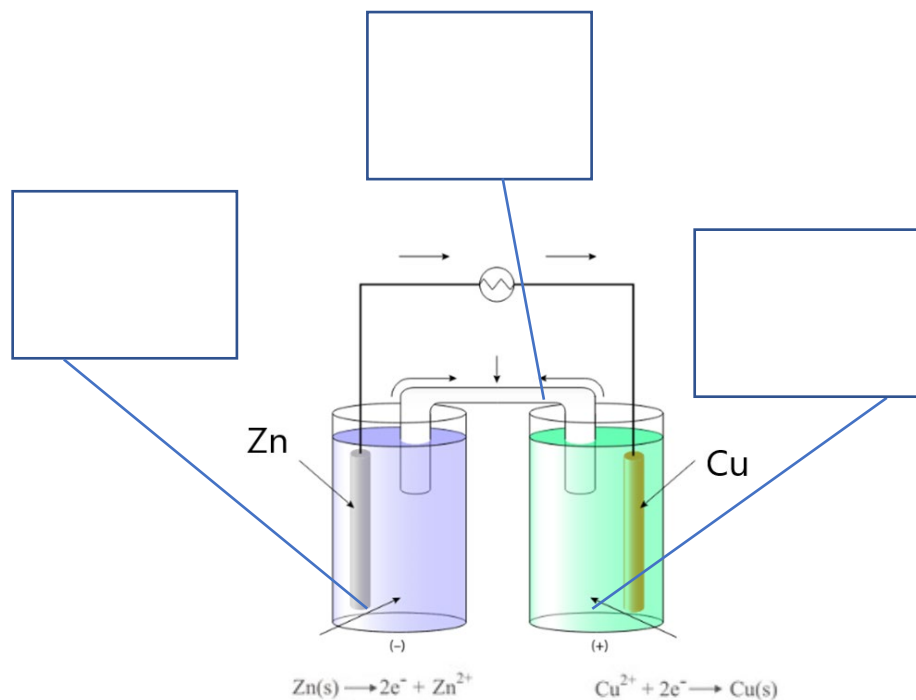


Half Reactions	E° (V)
$Zn^{2+}(aq) + 2e^-(aq) \rightarrow Zn(s)$	-0.76 V
$Cu^{2+}(aq) + 2e^-(aq) \rightarrow Cu(s)$	+0.34 V

Anode = oxidation reduction
 Anode Mass is: increasing decreasing staying the same

Cathode = oxidation reduction
 Cathode Mass is _____

Cell Potential =



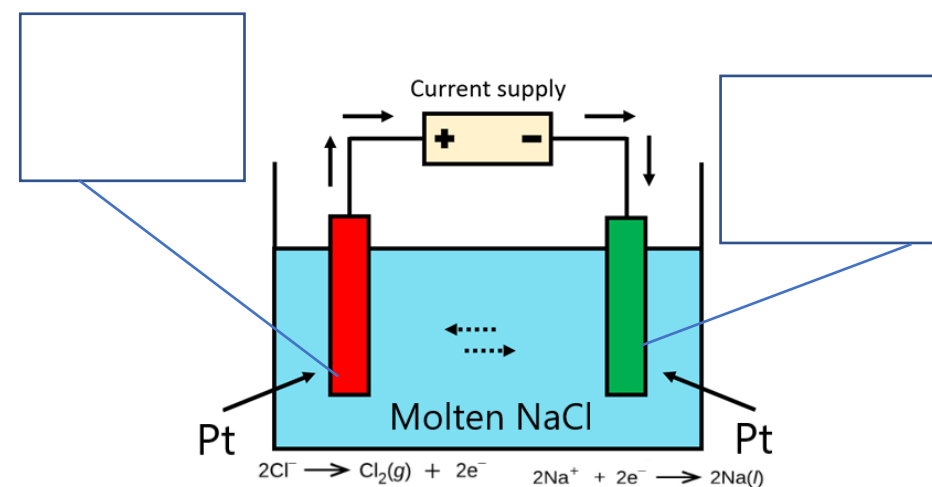
Free Energy Calculation:

Half Reactions	E° (V)
$Na^+(aq) + e^-(aq) \rightarrow Na(s)$	-2.71 V
$Cl_2(g) + 2e^-(aq) \rightarrow Cl^-(aq)$	+1.36 V

Anode = oxidation reduction
 Anode Mass is _____

Cathode = oxidation reduction
 Cathode Mass is _____

Cell Potential =



Free Energy Calculation:

Electrolytic Cells

Galvanic Cells