

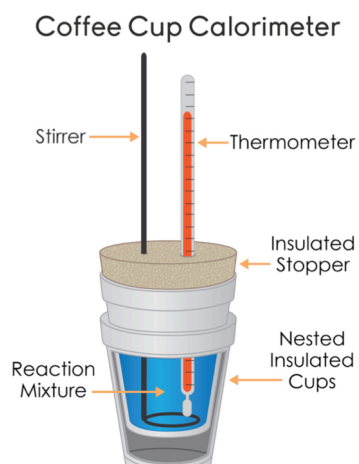
Units of Heat:

Cal / Joule

Heat Capacity

Specific Heat Capacity

Material	Specific Heat (Joules/gram·°C)
Liquid water	4.18
Solid Water (ice)	2.11
Wood (typical)	1.674
Air	1.020
Aluminum	0.900
Iron/Steel	0.444
Copper	0.387
Brass	0.380
Silver	0.240
Gold	0.130



$$Q = m c \Delta T$$

Calorimetry

Amount of Heat Required to: This is an _____ Property	Amount of Heat Required to: This is an _____ Property
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Applying It

Specific Heat Capacity: Unique for substances – can be used to ID substances	Measures _____ Absorbed or Released by a substance or reaction. KNOW THIS: All the Heat Entering the system Comes From: (and vice versa).
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Tying it Up in Calculations

Heat	=	Mass	Specific Heat	Temp Diff ($T_{\text{final}} - T_{\text{initial}}$)
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